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INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(51) International Patent Classification ⁶ :	A1	(11) International Publication Number: WO 97/03149
C10G 45/02		(43) International Publication Date: 30 January 1997 (30.01.97)
(21) International Application Number: PCT/US96/10971		(81) Designated States: AL, AM, AT, AU, AZ, BB, BG, BR, BY, CA, CH, CN, CZ, DE, DK, EE, ES, FI, GB, GE, HU, IS, JP, KE, KG, KP, KR, KZ, LK, LR, LS, LT, LU, LV, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, TJ, TM, TR, TT, UA, UG, UZ, VN, ARIPO patent (KE, LS, MW, SD, SZ, UG), Eurasian patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European patent (AT, BE, CH, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE), OAPI patent (BF, BJ, CF, CG, CI, CM, GA, GN, ML, MR, NE, SN, TD, TG).
(22) International Filing Date: 27 June 1996 (27.06.96)		
(30) Priority Data: 08/500,100 10 July 1995 (10.07.95) US		(Published)
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(54) Title: HYDRODESULFURIZATION PROCESS

(57) Abstract

A process for the hydrodesulfurization of petroleum streams is disclosed wherein the sulfur containing petroleum stream is contacted along with hydrogen at a partial pressure of less than 70 psig in a distillation column reactor containing a hydrodesulfurization catalyst in the form of a catalytic distillation structure.

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HYDRODESULFURIZATION PROCESS
BACKGROUND OF THE INVENTION

Field of the Invention

5 The present invention relates to the hydrodesulfurization of petroleum streams in a distillation column reactor. More particularly the invention relates to a process wherein a petroleum fraction is fed to a distillation column reactor containing a hydrodesulfurization catalyst in the form of a catalytic 10 distillation structure where the organic sulfur compounds contained in the petroleum fraction are reacted with hydrogen to form H₂S which can be stripped from the overhead product.

Related Information

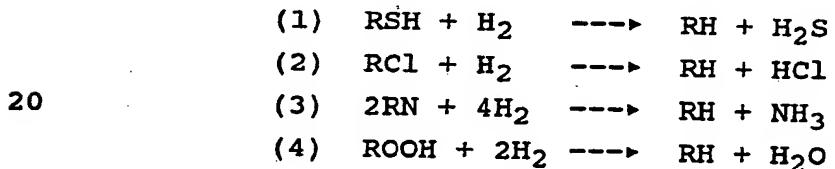
15 Petroleum distillate streams contain a variety of organic chemical components. Generally the streams are defined by their boiling ranges which determine the compositions. The processing of the streams also affects the composition. For instance, products from either catalytic cracking or thermal cracking processes contain 20 high concentrations of olefinic materials as well as saturated (alkanes) materials and polyunsaturated materials (diolefins). Additionally, these components may be any of the various isomers of the compounds.

25 Organic sulfur compounds present in these petroleum fractions are denoted as, "sulfur". The amount of sulfur is generally dependent upon the crude source. For instance the Saudi Arabian crudes are generally high in sulfur as are certain domestic crudes. Kuwaiti, Libyan and Louisiana 30 crudes are generally low in sulfur. The type of sulfur compounds will also depend on the boiling range of a given stream. Generally the lower boiling fractions contain mercaptans while the heavier boiling fractions contain thiophenic and heterocyclic sulfur compounds.

35 The organic sulfur compounds are almost always considered to be contaminants. They hinder in downstream processing and at the very least make obnoxious SO₂ gas when burned. For these reasons it is very desirable to

remove these compounds. The degree of removal is dependent upon the use of the fraction. For instance, feed streams to catalytic reforming require extremely low sulfur concentrations (less than 1 wppm). Current EPA regulations 5 call for combustible motor fuels such as gasoline, kerosene or diesel to have no more than about 500 wppm sulfur. It is expected that in the future the sulfur specification will be lowered to about 50 wppm, especially for gasoline.

10 The most common method of removal of the sulfur compounds is by hydrodesulfurization (HDS) in which the petroleum distillate is passed over a solid particulate catalyst comprising a hydrogenation metal supported on an alumina base. In the past this has generally been done by downflow over fixed beds concurrently with copious 15 quantities of hydrogen in the feed. The following equations illustrate the reactions in a prior art typical HDS unit:



Additional reactions depend upon the sulfur compounds present and the source of the fraction. The catalyst used for hydrodesulfurization necessarily is a hydrogenation 25 catalyst and the support sometimes is acidic in nature. The latter characteristics provide for some hydrocracking and hydrogenation of unsaturated compounds. The hydrocracking results in a higher volume of a less dense (lower specific gravity) material than the feed.

30 Typical operating conditions for the prior art HDS reactions are:

Temperature, °F	600-780
Pressure, psig	600-3000
H ₂ recycle rate, SCF/bbl	1500-3000
Fresh H ₂ makeup, SCF/bbl	700-1000

After the hydrotreating is complete the product is fractionated or simply flashed to release the hydrogen

sulfide and collect the now sweetened fraction.

It should be noted that the conditions or severity of the operation will depend upon the sulfur compounds present and the degree of desulfurization desired. For instance 5 mercaptans are much more easily desulfurized than thiophenes. The desulfurization of thiophenes and other heterocyclic sulfur compounds necessarily involves breaking and saturation of the rings which requires higher severity.

A method of carrying out catalytic reactions has been 10 developed wherein the components of the reaction system are concurrently separable by distillation using the catalyst structures as the distillation structures. Such systems are described variously in U.S. Pat. Nos. 4,215,011; 4,232,177; 4,242,530; 4,250,052; 4,302,356 and 4,307,254 15 commonly assigned herewith. In addition, commonly assigned U.S. Patent 4,443,559, 5,057,468, 5262,012 5,266,546 and 5,348,710 disclose a variety of catalyst structures for this use and are incorporated by reference herein. A 20 distillation column reactor has been utilized wherein a solid particulate catalyst has been placed within a distillation column so as to act as a distillation structure. The distillation column reactor has been found to be particularly useful in equilibrium limited reactions because the reaction products have been removed from the 25 reaction zone almost immediately. Additionally the distillation column reactor has been found to be useful to prevent unwanted side reactions.

In U.S. patent 4,194,964, Chen, et al propose a process 30 operated at about 300 psig to 3000 psig, high hydrogen partial pressures and high hydrogen flow rates (around 4000 SCF/B) for concurrent hydroprocessing and distillation of heavy petroleum stocks. Essentially Chen, et al disclose the use of concurrent distillation and hydroprocessing of the heavy stocks for the standard high pressure treating 35 and hydrocracking. The range of conditions is fairly consistent with the prior art processes.

SUMMARY OF THE INVENTION

The present invention uses catalytic distillation in

hydrodesulfurization at low total pressures below about 300 psig, preferably below about 290 and more preferably in the range of 0 to 200 psig, low hydrogen partial pressure in the range of 0.1 to 70 psi and temperatures in the 5 range of 400 to 800 °F. Briefly the invention may be said to comprise:

feeding (1) a petroleum stream containing sulfur compounds and (2) hydrogen to a distillation column reactor;

10 concurrently in said distillation column reactor

(a) distilling said petroleum stream whereby there are vaporous petroleum products rising upward through said distillation column reactor, an internal reflux of liquid flowing downward in said distillation column reactor and condensing petroleum products within said distillation column reactor,

15 (b) contacting said petroleum stream and said hydrogen in the presence of a hydrodesulfurization catalyst prepared in the form of a catalytic distillation structure at total pressure of less than about 300 psig, hydrogen partial pressure in the range of 0.1 to less than 70 psi and a temperature in the range of 400 to 800 °F whereby a portion of the sulfur compounds contained within said 20 petroleum stream react with hydrogen to form H_2S ;

25 withdrawing an overheads from said distillation column reactor containing said H_2S ;

separating the H_2S from said overheads by condensing a higher boiling fraction in a partial condenser;

30 returning a portion of said condensed overheads to said distillation column reactor as external reflux; and

35 withdrawing a bottoms product having a lower sulfur content than said petroleum stream.

DESCRIPTION OF THE PREFERRED EMBODIMENTS

The operation of the distillation column reactor results in both a liquid and vapor phase within the distillation reaction zone. A considerable portion of the vapor is hydrogen while a portion is vaporous hydrocarbon from the 35

petroleum fraction. Actual separation may only be a secondary consideration. Within the distillation reaction zone there is an internal reflux and liquid from an external reflux which cool the rising vaporous hydrocarbon condensing a portion within the bed.

5 Without limiting the scope of the invention it is proposed that the mechanism that produces the effectiveness of the present process is the condensation of a portion of the vapors in the reaction system, which occludes sufficient hydrogen in the condensed liquid to obtain the 10 requisite intimate contact between the hydrogen and the sulfur compounds in the presence of the catalyst to result in their hydrogenation.

15 The result of the operation of the process in the catalytic distillation mode is that lower hydrogen partial pressures (and thus lower total pressures) may be used.

20 As in any distillation there is a temperature gradient within the distillation column reactor. The temperature at the lower end of the column contains higher boiling material and thus is at a higher temperature than the upper end of the column. The lower boiling fraction, which 25 contains more easily removable sulfur compounds, is subjected to lower temperatures at the top of the column which provides for greater selectivity, that is, less hydrocracking or saturation of desirable olefinic compounds. The higher boiling portion is subjected to higher temperatures in the lower end of the distillation column reactor to crack open the sulfur containing ring compounds and hydrogenate the sulfur.

30 It is believed that in the present reaction catalytic distillation is a benefit first, because the reaction is occurring concurrently with distillation, the initial reaction products and other stream components are removed from the reaction zone as quickly as possible reducing the 35 likelihood of side reactions. Second, because all the components are boiling the temperature of reaction is controlled by the boiling point of the mixture at the system pressure. The heat of reaction simply creates more

boil up, but no increase in temperature at a given pressure. As a result, a great deal of control over the rate of reaction and distribution of products can be achieved by regulating the system pressure. Also, 5 adjusting the throughput (residence time = liquid hourly space velocity⁻¹) gives further control of product distribution and to a degree control of the side reactions such as oligomerization. A further benefit that this 10 reaction may gain from catalytic distillation is the washing effect that the internal reflux provides to the catalyst thereby reducing polymer build up and coking.

Finally, the upward flowing hydrogen acts as a stripping agent to help remove the H₂S which is produced in the distillation reaction zone.

15 Petroleum fractions which may be treated to remove sulfur by the instant process include the full range of petroleum distillates and include natural gas liquids, naphthas, kerosene, diesel, gas oils (both atmospheric and vacuum) and residuums. The fractions may be straight run 20 material direct from a crude fractionation unit or may be the result of downstream processing, such as fluid catalytic cracking, pyrolysis or delayed coking.

25 The hydrogen rate to the reactor must be sufficient to maintain the reaction, but kept below that which would cause flooding of the column which is understood to be the "effectuating amount of hydrogen" as that term is used herein. The mole ratio of hydrogen to sulfur compound in the feed varies according to the type of compound and the amount of hydrogen expected to be consumed by side reactions such as double or triple bond saturation. Hydrogen flow rates are typically calculated as standard cubic feet 30 per barrel of feed (SCFB) and are in the range of 300 to 3000 SCFB.

35 Surprisingly, a low total pressure, below about 300 psig, for example in the range of 0 to 200 psig is required for the hydrodesulfurization and hydrogen partial pressure of less than 70 psi down to 0.1 psig can be employed, e.g. 0.1 to 70 psig preferably about 0.5 to 10 psig. The

preferred hydrogen partial pressure is less than 50 psig. This preferably is a hydrogen partial pressure in the range of about 0.1 to 10 psia and even more preferably no more than 7 psia. Optimal results have been obtained in the range between 0.5 and 50 psig hydrogen partial pressure.

5 Where the petroleum stream is a naphtha, typical conditions are overhead temperature in the range of 350 to 550 °F, the bottoms temperature in the range of 500 to 800 °F, and the pressure in the range of 25 to less than 300 psig. Where the petroleum stream is a kerosene, typical conditions are overhead temperature in the range of 350 to 650 °F, the bottoms temperature in the range of 500 to 800 °F, and the pressure in the range of 0 to 200 psig. Where the petroleum stream is a diesel, typical conditions are overhead temperature in the range of 350 to 650 °F, the bottoms temperature in the range 500 to 850 °F, and the pressure in the range 0 to 150 psig.

10 15 Catalyst which are useful for the hydrodesulfurization reaction include metals Group VIII such as cobalt, nickel, palladium, alone or in combination with other metals such as molybdenum or tungsten on a suitable support which may be alumina, silica-alumina, titania-zirconia or the like. Normally the metals are provided as the oxides of the metals supported on extrudates or spheres and as such are not generally useful as distillation structures.

20 25 30 35 The catalysts contain components from Group V, VIB, VIII metals of the Periodic Table or mixtures thereof. The use of the distillation system reduces the deactivation and provides for longer runs than the fixed bed hydrogenation units of the prior art. The Group VIII metal provides increased overall average activity. Catalysts containing a Group VIB metal such as molybdenum and a Group VIII such as cobalt or nickel are preferred. Catalysts suitable for the hydrodesulfurization reaction include cobalt-molybdenum, nickel-molybdenum and nickel-tungsten. The metals are generally present as oxides supported on a neutral base such as alumina, silica-alumina or the like. The metals are reduced to the sulfide either in use or prior to use by

exposure to sulfur compound containing streams. The properties of a typical hydrodesulfurization catalyst in Table I below.

TABLE I

5	Manufacture	Criterion Catalyst Co.
	Designation	C-448
	Form	Tri-lobe Extrudate
	Nominal size	1.2 mm diameter
	Metal, Wt.5	
10	Cobalt	2-5%
	Molybdenum	5-20%
	Support	Alumina

15 Broadly stated, the catalytic material is a component of a distillation system functioning as both a catalyst and distillation packing, i.e., a packing for a distillation column having both a distillation function and a catalytic function.

20 The reaction system can be described as heterogenous since the catalyst remains a distinct entity.

25 A preferred catalyst structure for the present hydrogenation reaction comprises flexible, semi-rigid open mesh tubular material, such as stainless steel wire mesh, filled with a particulate catalytic material in one of several embodiments recently developed in conjunction with the present process.

30 One new catalyst structure developed for use in hydrogenations is described in US Pat. No. 5,266,546 which is incorporated herein in its entirety. Briefly the new catalyst structure is a catalytic distillation structure comprising flexible, semi-rigid open mesh tubular material, such as stainless steel wire mesh, filled with a particulate catalytic material said tubular material having two ends and having a length in the range of from about one-half to twice the diameter of said tubular material, a first end being sealed together along a first axis to form a first seam and a second end being sealed together along a second axis to form a second seam wherein the plane of the first seam along the axis of said tubular material and the

plane of the second seam along the axis of said tubular material bisect each other at an angle of about 15 to 90°.

US Patent No. 4,242,530 and US Pat. No. 4,443,559 which are incorporated herein, disclose supported catalyst in a plurality of pockets in a cloth belt or wire mesh tubular structures, which is supported in the distillation column reactor by open mesh knitted stainless steel wire by twisting the two together into a helix. U.S. Pat. No. 5,348,710, which is incorporated herein, describes several other suitable structures in the prior art and discloses new structures suitable for this process. Other catalytic distillation structures useful for this purpose are disclosed in U.S. patents 4,731,229 and 5,073,236 which are also incorporated by reference.

The particulate catalyst material may be a powder, small irregular chunks or fragments, small beads and the like. The particular form of the catalytic material in the structure is not critical, so long as sufficient surface area is provided to allow a reasonable reaction rate. The sizing of catalyst particles can be best determined for each catalytic material (since the porosity or available internal surface area will vary for different material and of course affect the activity of the catalytic material).

For the present hydrodesulfurizations the preferred catalyst structures for the packing are those employing the more open structure of permeable plates or screen wire.

EXAMPLES

In the examples 1-3 below a catalyst structure was prepared as disclosed in U.S. Pat No. 5,431,890, which is incorporated herein, containing 0.947 pounds of the catalytic material described in Table I and placed in the middle nineteen feet of a 20 foot tall 1 inch diameter distillation column reactor. There were 1/2 foot of inert packing in a rectifying section above the catalyst and 1/3 foot of inert packing in a stripping section below the catalyst. Liquid feed was fed to the distillation column reactor at either at about the mid point or below the catalyst bed and hydrogen was fed at the bottom of the

10

catalyst bed. In each of the examples there is a showing of a substantial reduction in the amount of organic sulfur in both the overheads and bottoms, the removed organic sulfur that has been converted to H_2S and separated 5 overhead by partial condensation of the overheads.

Example 1

A full boiling range naphtha was fed to the distillation column reactor containing a the catalyst prepared as noted above. Conditions and results are given in TABLE II below.

10

TABLE II

	Run No.	4	25HDS
	Hours	605.2	
	Feed		
15	rate, lbs/hr	1.00	
	total sulfur, wppm (mg)	925 (419)	
	H_2 -rate, SCFH (SCFB)	11.03 (3243)	
	Temperature, °F		
20	Overhead	364	
	Top Catalyst Bed	503	
	Mid Catalyst Bed	514	
	Bottom Catalyst Bed	580	
	Bottoms	679	
	Feed	401	
25	Total Pressure, psig	200	
	Hydrogen Partial pressure, psig	23.9	
30	Overhead		
	rate, lbs/hr	0.74	
	total sulfur, wppm (mg)	120 (40)	
35	Bottoms		
	rate, lbs/hr	0.20	
	total sulfur, wppm (mg)	203 (18)	
	Reflux Ratio	10:1	
40	Catalyst above feed, feet	9	
	Catalyst below feed, feet	10	
	Conversion of organic S, %	86	

45

Example 2

A kerosene fraction was fed to the distillation column reactor described above. Conditions and results are given in TABLE III below.

11

TABLE III

Run No.	4	25HDS
Hours	1757.2	
5	Feed	
	rate, lbs/hr	1.00
	total sulfur, wppm (mg)	1528 (694)
	H ₂ rate, SCFH (SCFB)	5.02 (1476)
10	Temperature, °F	
	Overheads	449
	Top Catalyst Bed	647
	Mid Catalyst Bed	659
	Bottom Catalyst Bed	697
15	Bottoms	784
	Feed	450
	Total Pressure, psig	100
	Hydrogen Partial Pressure, psig	11
20	Overhead	
	rate, lbs/hr	0.81
	total sulfur, wppm (mg)	38 (14)
25	Bottoms	
	rate, lbs/hr	0.17
	total sulfur, wppm (mg)	1577 (122)
30	Reflux Ratio	10:1
	Catalyst above feed, feet	19
	Catalyst below feed, feet	0
35	Conversion of organic S, %	80

Example 3

40 A diesel fraction was fed to the distillation column reactor as described above. Conditions and results are given in TABLE IV below.

TABLE IV

Run No.	4	25HDS
Hours	1421.2	
5	Feed	
	rate, lbs/hr	1.00
	total sulfur, wppm (mg)	1528 (694)
	H ₂ rate, SCFH (SCFB)	10.03 (2949)
10	Temperature, °F	
	Overheads	438
	Top Catalyst Bed	634
	Mid Catalyst Bed	648
	Bottom Catalyst Bed	689
15	Bottoms	801
	Feed	450
	Total Pressure, psig	100
	Hydrogen Partial Pressure, psig	23
20	Overhead	
	rate, lbs/hr	0.77
	total sulfur, wppm (mg)	84 (29)
25	Bottoms	
	rate, lbs/hr	0.21
	total sulfur, wppm (mg)	1278 (122)
30	Reflux Ratio	10:1
	Catalyst above feed, feet	19
	Catalyst below feed, feet	0
35	Conversion of organic S, %	78

Example 4

In the following example 18.7 pounds of the catalytic material of Table I were placed in the catalytic distillation structure prepared as disclosed in U.S. Pat. No. 5,431,890, and were placed in the mid 40 feet of a fifty foot tall three inch distillation column reactor. Liquid feed was to about two thirds of the way up the column and hydrogen was fed below the bed. A second full range fluid cracked naphtha was feed to the column and the conditions and results are reported in TABLE V below.

13

TABLE V

Run No.	1027HDS
Hours	308
5 Feed	
rate, lbs/hr	20.0
total sulfur, wppm (mg)	1242 (11,277)
H ₂ rate, SCFH (SCFB)	41 (601)
10 Temperature, °F	
Overheads	476
Top Catalyst Bed	552
Mid Catalyst Bed	651
Bottom Catalyst Bed	696
15 Bottoms	749
Feed	297
Total Pressure, psig	200
Hydrogen Partial Pressure, psig	21.23
20 Overhead	
rate, lbs/hr	16.0
total sulfur, wppm (mg)	122 (886)
25 Bottoms	
rate, lbs/hr	4.0
total sulfur, wppm (mg)	35 (64)
30 Reflux Ratio	4:1
Internal Reflux	6.02
Catalyst above feed, feet	15
Catalyst below feed, feet	25
35 Conversion of organic S, %	92

The invention claimed is:

1. A process for the hydrodesulfurization of petroleum streams comprising:

(A) feeding (1) a petroleum stream containing organic sulfur compounds and (2) hydrogen to a distillation column reactor;

(B) concurrently in said distillation column reactor

(1) distilling said petroleum stream whereby there are vaporous petroleum products rising upward through said distillation column reactor, an internal reflux of liquid flowing downward in said distillation column reactor and condensing petroleum products within said distillation column reactor, and

(2) contacting said petroleum stream and said hydrogen in the presence of a hydrodesulfurization catalyst prepared in the form of a catalytic distillation structure at total pressure less than about 300 psig, hydrogen partial pressure in the range of 0.01 to less than 70 psi and a temperature in the range of 400 to 800 °F whereby a portion of the organic sulfur compounds contained within said petroleum stream react with hydrogen to form H₂S;

(C) withdrawing an overheads from said distillation column reactor containing said H₂S;

(D) separating the H₂S from said overheads by condensing a higher boiling fraction;

(E) returning a portion of said condensed overheads to said distillation column reactor as external reflux; and

(F) withdrawing a bottoms product having a lower sulfur content than said petroleum stream.

2. The process according to claim 1 wherein the pressure is below about 290 psig.

3. The process according to claim 1 wherein the pressure is in the range of 0 to 200 psig.

4. The process according to claim 1 wherein the hydrogen partial pressure is in the range of about 0.1 to 70 psig.

5. The process according to claim 1 wherein the H₂S is separated from said overheads by condensing the higher

boiling fraction in a partial condenser and removing said H₂S as a gas along with any unreacted hydrogen.

6. The process according to claim 5 wherein a portion of said condensed overheads is returned to said distillation column reactor as external reflux.

7. The process according to claim 1 wherein said petroleum stream is a naphtha.

8. The process according to claim 7 wherein the overhead temperature is in the range of 350 to 550 °F, the bottoms temperature is in the range of 500 to 800 °F, and the pressure is in the range of 25 to 300 psig.

9. The process according to claim 1 wherein said petroleum stream is a kerosene.

10. The process according to claim 9 wherein the overhead temperature is in the range of 350 to 650 °F, the bottoms temperature is in the range of 500 to 800 °F, and the pressure is in the range of 0 to 200 psig.

11. The process according to claim 1 wherein said petroleum stream is a diesel.

12. The process according to claim 11 wherein the overhead temperature is in the range of 350 to 650 °F, the bottoms temperature is in the range 500 to 850 °F, and the pressure is in the range 0 to 150 psig.

13. The process according to claim 6 wherein all of said condensed overheads is returned to said distillation column reactor as external reflux.

14. The process according to claim 1 wherein a bottoms product is taken from said distillation column reactor.

15. The process according to claim 14 wherein a portion of said bottoms is heated in a reboiler and returned to said distillation column reactor.

16. The process according to claim 14 wherein all of said bottoms is heated in a reboiler and returned to said distillation column reactor.

17. The process according to claim 1 wherein the H₂S is removed from said overheads by stripping in a separate distillation column.

18. The process according to claim 1 wherein the

hydrogen partial pressure is in the range of about 0.1 to 70 psi.

INTERNATIONAL SEARCH REPORT

International application No.

PCT/US96/10971

A. CLASSIFICATION OF SUBJECT MATTER

IPC(6) :C10G 45/02

US CL :208/213, 216R, 217; 203/28, 29

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

U.S. : 208/213, 216R, 217; 203/28, 29

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)

APS-search terms: reactive distillation

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
Y	US, A, 4,194,964 (Chen et al.) 25 March 1980, col. 3, line 46 through col. 4, line 18; col. 4, lines 34-49; col. 4, line 58 through col. 6, line 6, col. 8, lines 3-19; abstract.	1-18
Y	Hobson, G.D., "Modern Petroleum Technology", 06 June 1973, pages 363-371, especially pages 364-366.	1-18
A	US, A, 4,213,847 (CHEN ET AL.) 22 July 1980.	1-18
A	US, A, 4,232,177 (SMITH, JR.) 04 November 1980.	1-18
A	US, A, 5,173,173 (VAUK) 22 December 1992.	1-18

<input type="checkbox"/>	Further documents are listed in the continuation of Box C.	<input type="checkbox"/>	See patent family annex.
•	Special categories of cited documents:	"T"	later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention
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13 AUGUST 1996		29 NOV 1996	
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